
	INDIAN SCHOOL AL WADI AL KABIR		
Class: XI	DEPARTMENT: SCIENCE 2022-23 SUBJECT: CHEMISTRY		Date of completion: II week of November 2022
Worksheet No: 05 with answers	TOPIC: REDOX REACTIONS		Note: A4 FILE FORMAT
NAME OF THE STUDENT	CLASS & SEC:	ROLL NO.	

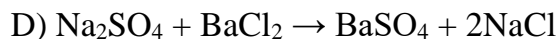
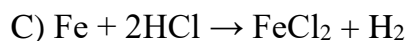
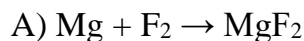
MULTIPLE CHOICE QUESTIONS

- Using Stock notation, represent the compound MnCl_4
 A) Mn(II)Cl_4 B) Mn(III)Cl_4 C) Mn(IV)Cl_4 D) Mn(VI)Cl_4
- What is the oxidation number of Mn in K_2MnO_4 ?
 A) +5 B) +6 C) +7 D) +3
- In which compound is the oxidation number of nitrogen +5?
 A) $\text{Ca(NO}_3)_2$ B) Mg_3N_2 C) N_2O_3 D) NaNO_2
- In which reaction is Cu^{2+} acting as an oxidising agent?
 A) $\text{Cu}^{2+} + 2\text{Ag} \rightarrow 2\text{Ag}^+ + \text{Cu}$ B) $2\text{Cu}^+ + \text{O}^{2-} \rightarrow \text{Cu}_2\text{O}$
 C) $3\text{Cu} + \text{O}_2 \rightarrow \text{Cu}_2\text{O} + \text{CuO}$ D) $\text{Cu} + \text{Hg}^{2+} \rightarrow \text{Hg} + \text{Cu}^{2+}$
- Which statement about the following reaction is correct?

$$\text{Mg(s)} + 2\text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$$
 A) magnesium acts as oxidising agent
 B) hydrogen molecules act as reducing agent
 C) hydrogen ion acts as oxidising agent

D) chloride ion acts as oxidising agent

6. Which of the following is not a redox reaction?



7. Identify the element which undergoes disproportionation in the following reaction.



A) Na B) Cl

C) O D) H

Assertion Reason Type

A) Both Assertion and Reason are correct statements and Reason is the correct explanation of the Assertion.

B) Both Assertion and Reason are correct statements and Reason is not the correct explanation of the Assertion.

C) Assertion is correct but Reason is wrong statement.

D) Assertion is wrong but Reason is correct statement.

8. **Assertion (A):** The decomposition of hydrogen peroxide to form water and oxygen is an example of disproportionation reaction.

Reason (R): The oxygen of peroxide is in -1 oxidation state and it is converted to zero oxidation state in O_2 and -2 in H_2O .

9. **Assertion (A):** Hydrogen reacts with oxygen to form water is a redox reaction.

Reason (R): Hydrogen acts as an oxidising agent and oxygen as reducing agent.

10. **Assertion (A):** Standard electrode potential of potassium is more positive than that of Aluminium.

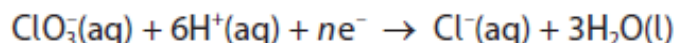
Reason (R): Potassium is a stronger reducing agent than aluminium.

Read the given passage and answer the questions that follow:

Redox reactions are those reactions in which oxidation and reduction occur simultaneously. A redox reaction is made up of two half reactions. In the first half reaction, oxidation takes place and second half reduction occurs. Oxidation is a process in which a substance loses electrons and in reduction substance gains electrons. The substance which gains electrons is reduced and acts as an oxidizing agent. On the other hand, a substance which loses electrons is oxidized and acts as a reducing agent. The oxidation number of an atom increases during oxidation and reduces during reduction. The redox reactions may include combination of atoms or molecules displacement of metals or non-metals and disproportionation reaction.

11. What is the oxidation number of oxygen in O_2F_2 ? (1)

12. What is the value of n in the following ionic half equation? (1)



13. Write formulas for Tin(IV) oxide. (1)

14. In which of the following compounds does oxygen show a positive oxidation state?



15. ClO_4^- does not undergo disproportionation. Why? (1)

Question – Answer Type:

16. What are Disproportionation reactions? Give an example. (1)

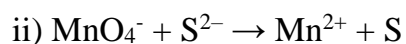
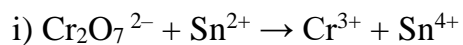
17. Using Stock notation, represent the following compounds: (1)
i) H_2CrO_4 ii) Cu_2O

18. Given the standard electrode potentials, (1)
 $K^+/K = -2.93V$, $Ag^+/Ag = 0.80V$, $Ni^{2+}/Ni = -0.25V$,
 $Ca^{2+}/Ca = -2.87V$, $Al^{3+}/Al = -1.66V$. Arrange these metals in their increasing order of reducing power.

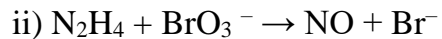
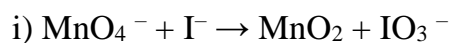
19. Define a redox couple. (1)

20. $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$ is not a redox reaction. Why? (1)

21. Balance the following redox reactions in acidic medium. (2 x 2)



22. Balance the following redox reactions in basic medium. (2 x 2)



ANSWERS

1.	C) Mn(IV)Cl_4
2.	B) +6
3.	A) $\text{Ca(NO}_3)_2$
4.	A) $\text{Cu}^{2+} + 2\text{Ag} \rightarrow 2\text{Ag}^+ + \text{Cu}$
5.	C) hydrogen ions act as oxidising agents
6.	D) $\text{Na}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$
7.	B) Cl

8.	A) Both Assertion and Reason are correct statements and Reason is the correct explanation of the Assertion
9.	C) Assertion is correct but Reason is wrong statement.
10.	D) Assertion is wrong but Reason is correct statement.
11.	+1
12.	n = 6
13.	SnO ₂
14.	OF ₂
15.	In ClO ₄ ⁻ , chlorine is present in its highest oxidation state.
16.	Disproportionation reactions are redox reactions in which an element in one oxidation state is simultaneously oxidised and reduced. Eg:- 2H ₂ O ₂ (aq) → 2H ₂ O(l) + O ₂ (g)
17.	i) H ₂ Cr(VI)O ₄ ii) Cu ₂ (I)O
18.	Ag ⁺ /Ag < Ni ²⁺ /Ni < Al ³⁺ /Al < Ca ²⁺ /Ca < K ⁺ /K
19.	A redox couple is defined as having together the oxidised and reduced forms of a substance taking part in an oxidation or reduction half reaction.
20.	No change in oxidation number for any of the elements.
21.	i) Cr ₂ O ₇ ²⁻ + 14H ⁺ + 3 Sn ²⁺ → 2Cr ³⁺ + 3Sn ⁴⁺ + 7H ₂ O ii) 2MnO ₄ ⁻ + 5S ²⁻ + 16H ⁺ → 2Mn ²⁺ + 8H ₂ O + 5S
22.	i) 2MnO ₄ ⁻ + H ₂ O + I ⁻ → 2MnO ₂ + 2OH ⁻ + IO ₃ ⁻ ii) 3N ₂ H ₄ + 4BrO ₃ ⁻ → 6NO + 4Br ⁻ + 6H ₂ O
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